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Research Article

French Fries-Like Bismuth Oxide: Physicochemical Properties, Electrical Conductivity and Photocatalytic Activity

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Abstract

Bismuth oxide synthesis using hydrothermal method has been conducted. This study aims to examine the effect of the hydrothermal reaction time on product characteristics and photocatalytic activity in degrading methyl orange dye. Bismuth oxide synthesis was initiated by dissolving bismuth nitrate pentahydrate ($\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$) and Na_2SO_4 in distilled water and added NaOH gradually. The solution formed was transferred into a Teflon-lined autoclave and heated at 120°C with time variations of 8–16 h. The formation of bismuth oxide was indicated by the vibrations of the Bi–O–Bi and Bi–O groups and the crystal structure consisting of $\alpha\text{-Bi}_2\text{O}_3$, $\beta\text{-Bi}_2\text{O}_3$, and $\gamma\text{-Bi}_2\text{O}_3$. In addition, the highest photocatalytic activity can be examined through several factors, such as: content of Bi–O–Bi and Bi–OH groups, crystal structure, band gap values, morphology, and surface area, acquired as a result of the effect of hydrothermal reaction time.

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Keywords: Bismuth Oxide; Bi_2O_3 ; Semiconductor; Hydrothermal; Photocatalysis

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1. Introduction

Bismuth oxide exhibits several useful properties, two of which are photocatalytic and electrochemical activities. Photocatalysis combines photochemical processes and catalysis in one framework, that is, chemical degradation processes that involve light as a trigger and a cata-

lyst as an accelerator of the transformation processes therein [1]. Meanwhile, bismuth oxide is also found to be potential as battery anode which could be attributed to large specific capacity, good electrical and optical properties with a bandgap of 2.8 eV, and a potential difference of 1.75-2.25 V [2]. Both potential properties are influenced by physicochemical characteristics of materials, such as crystal structure, morphology, functional group, band gap, which depend on material fabrication [3–6].

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Some of the experimental investigations and preparation processes have demonstrated the strong correlation between the synthesis procedure of Bi_2O_3 and its photo-catalytic activities [7]. The different methods and experimental conditions of synthesis result in production of different size, morphology, crystal structure and band gap of the resulting materials affecting the material's photocatalytic activity [8,9] and other properties, such as electrochemical behavior [10].

Synthesis of bismuth oxide can be conducted by several methods including precipitation [5,11], sol-gel [6], solution combustion [4,12,13] and hydrothermal [14–17]. In this study, the hydrothermal method was applied to produce bismuth oxide since it has several advantages, such as: precipitant-free, additive-free, no need sophisticated equipment, and low temperature processing, which lead to energy saving and cost effective [14–18]. Moreover, controllable particle size, morphology and the degree of crystallinity of the resulting products can be achieved [14]. The factors that affect the synthesized products in this method include temperature, pressure, and hydrothermal process time. The heat produced by water in the hydrothermal method provides a vapor pressure that plays a role in crystal formation. Several studies have studied bismuth oxide synthesized by hydrothermal method [14–19]. However, the overall properties of the resulting products, such as: crystal structure, morphology, purity, band gap, electrical properties and photocatalytic activity, have not been fully studied comprehensively. Taking these accounts, the present research delves into the synthesis of bismuth oxide using the hydrothermal method and investigate physicochemical properties, electrical conductivity, and photocatalytic activity.

Bismuth oxide is a potential material for a photocatalyst and has been applied to degrade dye molecules, such as methyl orange, remazol black B (RBB) and rhodamine B (RhB) [4,13,20,21]. Industries, such as printing, textiles, cosmetics and papers, are inseparable from the use of dyes, especially synthetic dyes [22,23]. These dyes have persistent stability to chemical substances, heat and light and are endowed with prolonged endurance in the environment. Because of their nature, the industrial waste produced is practically harmful to aquatic environments and interferes with light penetration in the receiving water bodies resulting in disturbing the biological processes [22,23]. Therefore, it is necessary to remove toxic dyes before discharging them into the environment

which can be undertaken by oxidation processes (AOPs) based on the oxidation through reaction with hydroxyl radicals ($\cdot\text{OH}$). Photocatalysis has been established as an effective advanced oxidation process [24] and of great interest due to efficiency, environmentally-friendly and high-throughput properties [25].

2. Materials and Methods

2.1 Materials

The materials used in this study were bismuth nitrate pentahydrate ($\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$) from Sigma-Aldrich. Other reagents used were sodium sulfate (Na_2SO_4), sodium hydroxide (NaOH), and methyl orange bought from Merck, Indonesia. Moreover, distilled water was also used.

2.2 Bismuth Oxide Synthesis

The synthesis of bismuth oxide was done using the hydrothermal method. The synthesis was initiated by adding 40 mL of distilled water into a beaker containing 8 mmol $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ and 12 mmol Na_2SO_4 , which was then stirred using a magnetic stirrer for 45 minutes. Subsequently, 72 mmol of NaOH was dissolved in 40 mL of distilled water and was added gradually to the previous solution. Once homogenous, the mixture was transferred into the Teflon-lined autoclave (a hydrothermal reactor) and heated at a temperature of 136°C with time variations of 8, 10, 12, 14, and 16 h. The mixture was then cooled to room temperature. The precipitate obtained was washed using distilled water, then dried at 80°C .

2.3 Material Characterizations

The powders obtained from the synthesis were characterized using X-ray diffraction (XRD), Fourier transform - infra red (FT-IR), scanning electron microscopy (SEM), and UV-vis diffuse reflectance spectroscopy (DRS) instruments. The crystal structure of the bismuth oxide synthesized was analyzed using the XRD Shimadzu 7000. Infrared spectroscopic analysis was carried out using the Shimadzu IRAffinity-1 FTIR in the wavenumber range of $500\text{--}4500\text{ cm}^{-1}$ to determine the functional groups present in the synthesized products. The morphology of the synthesized bismuth oxide was seen using SEM JEOL JED 6510LA with magnifications of 1000 and 5000 times. Pharmaspec UV-DRS UV 1700 analysis was performed to calculate the bandgap of the synthesized bismuth oxide. These characteriza-

tions also aimed to identify the purity of the obtained products. Moreover, electrical conductivity property was identified using LCR meter.

2.4 Photocatalytic Activity Test

In this photocatalytic test, a methyl orange stock solution of 1000 ppm was first made by dissolving 0.1 gram of methyl orange in 100 mL of distilled water. From that, methyl orange solution was made for photocatalysis. This solution was analyzed using a UV-Vis spectrophotometer to determine the maximum wavelength of the methyl orange. A series of standard solutions of 6 ppm, 5 ppm, 4 ppm, 3 ppm, and 2 ppm were made and measured for their absorbance with a UV-Vis spectrophotometer at a wavelength of 463 nm to obtain standard curve of methyl orange absorption. The photocatalytic process was carried out in a photocatalysis reactor. A total of 0.1 gram of the synthesized Bi_2O_3 was added into 50 mL of MO solution. The mixture was then stirred using a magnetic stirrer at a speed of 667 rpm. The stirring time was 2 h for the photocatalytic test without light. Meanwhile, for the photocatalytic test with light, time variations of 2, 4, 6, 8, and 10 h were used. Light irradiation was carried out using a UV-A lamp (315–400 nm) with a power of 15 watts. The final step was measurement of the methyl orange absorbance of the solutions that had been treated without and 31 h radiation at varying time duration using a UV-Vis spectrophotometer at a wavelength of 463 nm.

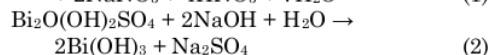
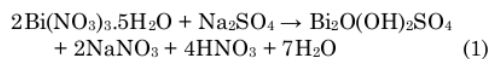
3. Results and Discussion

3.1 Synthesized Bismuth Oxide

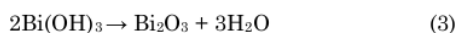
Bismuth oxide was synthesized using the hydrothermal method with time variations. The experiment aims to determine the effect of variations in hydrothermal time on the characteristics of the synthesized product. The hydrothermal time is said to affect the purity, crystal system, morphology, and crystal size, as well as the bandgap of the resulting product [26].

The synthesis was initiated by dissolving bismuth nitrate pentahydrate ($\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$) and sodium sulfate (Na_2SO_4) in distilled water then stirring them until homogenous. Na_2SO_4 would react with $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ to form $\text{Bi}_2\text{O}(\text{OH})_2\text{SO}_4$ complex, generating a product smaller in size than its precursors [14]. From the reaction of $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ and Na_2SO_4 , a white suspension was obtained. After the formation of the white suspension, NaOH was added. The OH^- ion from NaOH would react with $\text{Bi}_2\text{O}(\text{OH})_2\text{SO}_4$ to form a white precipitate

of $\text{Bi}(\text{OH})_3$. The reaction mechanism that occurred is as follows:



The white precipitate was then transferred into the Teflon-lined autoclave and then into the hydrothermal reactor to be heated at a temperature of 120 °C with time variations of 8, 10, 12, 14, and 16 h. When $\text{Bi}(\text{OH})_3$ is heated in a hydrothermal reactor, the compound will dehydrate and turn into Bi_2O_3 [14]. Subsequently, the obtained product was dried using an oven to remove water molecules that were still present in the Bi_2O_3 solid. The reaction mechanism that occurred in the hydrothermal reactor is as follows [14].



The five products obtained were in the form of powder with the same pale-yellow color. The pale yellow color indicated the presence of bismuth oxide in the synthesized products as it is one of the physical characteristics of bismuth oxide. The percentages of the yield of bismuth oxide with the hydrothermal times of 8, 10, 12, 14, and 16 h were 81.7%, 84.4%, 87.6%, 85.1%, and 82.1%, respectively.

3.2 Physicochemical and Electrical Properties

The XRD diffractograms of the bismuth oxide with variations in heating time are shown

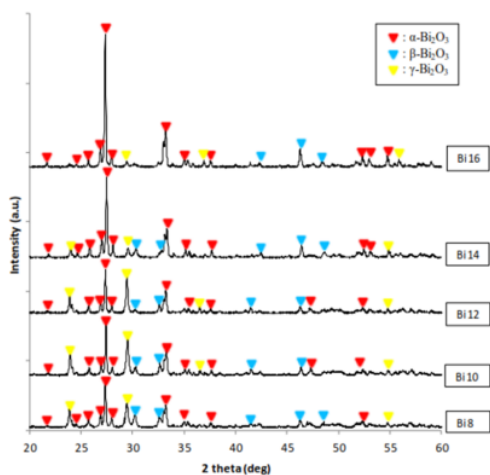


Figure 1. XRD diffractogram of bismuth oxide with variation in hydrothermal time.

in Figure 1. XRD data interpretation was undertaken by fitting the Bi_2O_3 diffractogram peaks using the Origin Pro 8 software to identify the 2θ value precisely. The fitting data were then compared with the joint committee database on powder diffraction standards (JCPDS) to determine the sample crystal system.

Bismuth oxides with hydrothermal times of 8, 10, 12, 14, and 16 h possessed the same mixed crystal structure, namely $\alpha\text{-Bi}_2\text{O}_3$ (monoclinic), $\beta\text{-Bi}_2\text{O}_3$ (tetragonal), and $\gamma\text{-Bi}_2\text{O}_3$ (cubic). The 2θ data of the bismuth oxide synthesized are shown in Table 1. The data corresponded to the JCPDS database number 41-1449 for $\alpha\text{-Bi}_2\text{O}_3$, the JCPDS database number 27-0050 for $\beta\text{-Bi}_2\text{O}_3$, and the JCPDS database number 45-1344 for $\gamma\text{-Bi}_2\text{O}_3$.

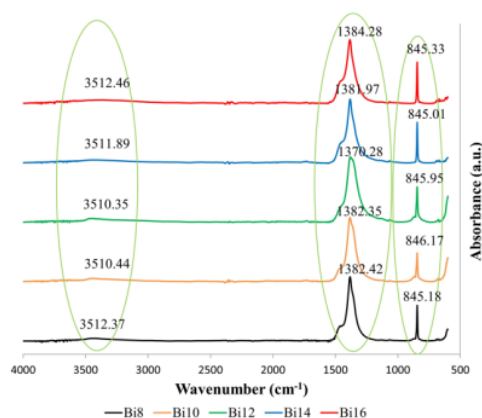


Figure 2. FTIR spectra of the synthesized bismuth oxide of varying hydrothermal times

As seen in Figure 1, the peak intensity ratios of $\alpha\text{-Bi}_2\text{O}_3$ to $\gamma\text{-Bi}_2\text{O}_3$ of the five samples were all different. The highest peak intensity ratio of $\alpha\text{-Bi}_2\text{O}_3$ to $\gamma\text{-Bi}_2\text{O}_3$ was found in Bi_2O_3 (16 h), while the lowest was in Bi_2O_3 (12 h). The higher peak ratio of $\alpha\text{-Bi}_2\text{O}_3$ to $\gamma\text{-Bi}_2\text{O}_3$ indicates that the intensity of the monoclinic crystal system ($\alpha\text{-Bi}_2\text{O}_3$) contained in a sample is higher, while the intensity of the cubic crystal system ($\gamma\text{-Bi}_2\text{O}_3$) is lower. This implied that Bi_2O_3 (16 h) had the highest number of monoclinic crystal systems ($\alpha\text{-Bi}_2\text{O}_3$) compared to the other hydrothermal time variations. Alternatively, Bi_2O_3 (12 h) had the lowest number of monoclinic crystal systems, though, this also indicated that Bi_2O_3 (12 h) had the highest number of cubic crystalline system ($\gamma\text{-Bi}_2\text{O}_3$) compared to other Bi_2O_3 . In the five-bismuth oxide diffractograms, it was also seen that the diffractogram peaks matched most with the JCPDS database of $\alpha\text{-Bi}_2\text{O}_3$ indicating that in the five bismuth oxide samples synthesized, the crystal system that was most abundant in the sample was the monoclinic crystal system of $\alpha\text{-Bi}_2\text{O}_3$.

From FTIR analysis, the five Bi_2O_3 samples obtained absorption peaks of the same functional groups though with different intensities. The FTIR spectra of the synthesized Bi_2O_3 are shown in Figure 2. The presence of absorptions at the wavenumber 845 cm^{-1} indicated the presence of a Bi-O-Bi group in the product [27] and at 1380 cm^{-1} , the presence of Bi-OH group in the product [28]. This is evidenced by the finding of a small absorption at wavenumber 3500 cm^{-1} which indicates the presence of the OH group [28].

Table 1. XRD 2θ data of the synthesized bismuth oxides.

Sample Code	Crystal Phase		
	$\alpha\text{-Bi}_2\text{O}_3$	$\beta\text{-Bi}_2\text{O}_3$	$\gamma\text{-Bi}_2\text{O}_3$
	2θ (°)	2θ (°)	2θ (°)
Bi8	26.88	30.22	23.87
	27.35	32.66	29.44
	33.21	46.29	54.77
Bi10	26.95	30.25	23.94
	27.42	32.65	29.51
	33.28	46.34	36.57
Bi12	26.89	32.58	23.88
	27.37	41.48	29.45
	33.23	46.29	36.50
Bi14	27.01	30.32	23.99
	27.48	42.45	29.56
	33.33	46.42	54.87
Bi16	26.88	42.32	29.43
	27.34	46.29	36.92
	33.21	48.83	54.74

The wavenumber range of 1250–1500 cm^{-1} , as presented in Figure 2, was assumed to represent several groups, such as NO_3^- , $\text{N}=\text{O}$, $\text{O}=\text{S}=\text{O}$, and $\text{Bi}-\text{OH}$, that concurrently absorbed the infrared radiation resulting in a wide and sharp absorption peak. These functional groups were theorized to have come from the by-products generated during the synthesis

process. With due regard, deconvolution peaks were made with the *fityk* application using the Gaussian function. The deconvolution peaks of Bi8, Bi10, Bi12, Bi14 and Bi16 are shown in Figure 3.

Figure 3 shows the four peak components that formed the peak that had been deconvoluted. The peak component around 1409 cm^{-1}

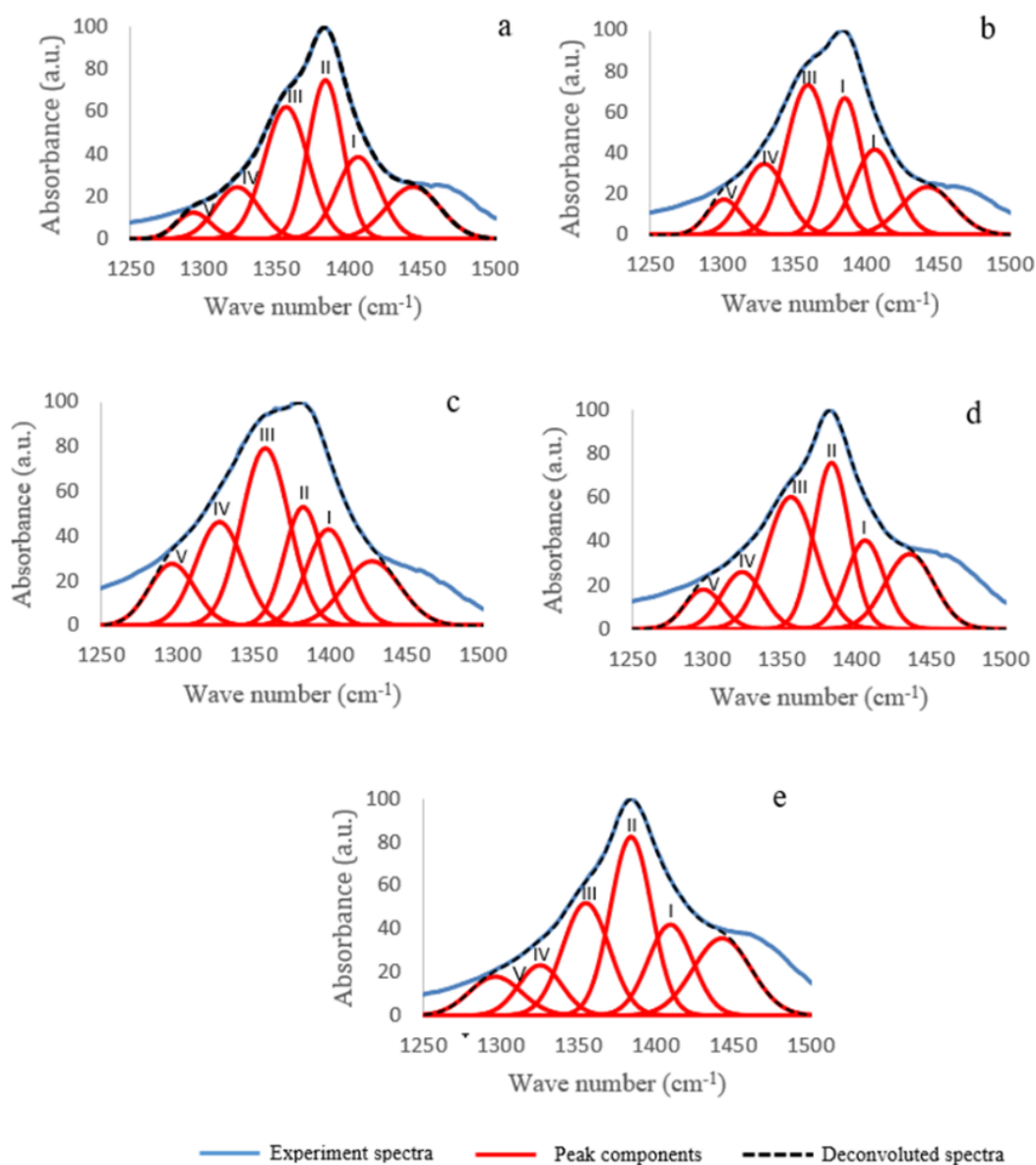


Figure 3. Deconvolution of FT-IR data of the synthesized bismuth oxide (a) Bi8; (b) Bi10; (c) Bi12; (d) Bi14; (e) Bi16.

(I) indicated the stretching vibration of the N=O group. Moreover, the component around 1325 cm^{-1} (IV) came from the asymmetric stretching vibration absorption of O=S=O, whereas the wavenumber 1295 cm^{-1} (V) was from NO_3^- asymmetric stretching vibration absorption [29]. The peak component around 1380 cm^{-1} (II) was most likely the stretching vibration of the Bi-OH group [28].

The presence of several functional groups such as NO_3^- , N=O dan O=S=O might have originated from the by-products generated during the synthesis process. The peak component II, which indicated the presence of the Bi-OH group, had the lowest intensity in bismuth oxide with a hydrothermal time of 12 h. The Bi-OH group most likely came from the $\text{Bi}(\text{OH})_3$ compound formed after the addition of NaOH during the synthesis process. These facts indicate that the 12 h sample had the highest $\text{Bi}(\text{OH})_3$ dehydration into Bi_2O_3 among other heating times.

In addition, the presence of peak component III at the wavenumber around 1380 cm^{-1} was a strong indication of the presence of the Bi-O-Bi group as evidenced by the FT-IR analysis of pure Bi_2O_3 from Sigma-Aldrich [13]. The

highest intensity of peak component III was found in Bi_2O_3 with a hydrothermal time of 12 h.

The SEM analysis was carried out on products with the highest and lowest photocatalytic activity (see section 3.26) namely Bi12 and Bi16. The SEM images are shown in Figure 4. SEM images show that Bi12 and Bi16 have almost the same French fries-like shape. The lengths of the Bi12 and Bi16 were $5.3\text{--}8.3\text{ }\mu\text{m}$ and $7.1\text{--}10.3\text{ }\mu\text{m}$ and their widths were $0.6\text{--}1.2$

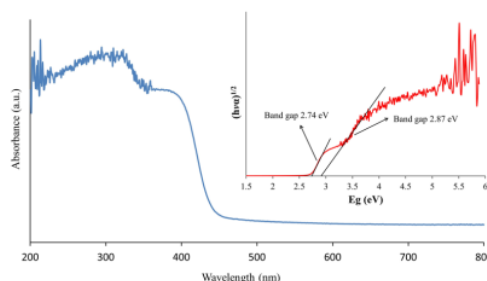


Figure 5. UV-DRS spectra of Bi_2O_3 with 12 h hydrothermal time.

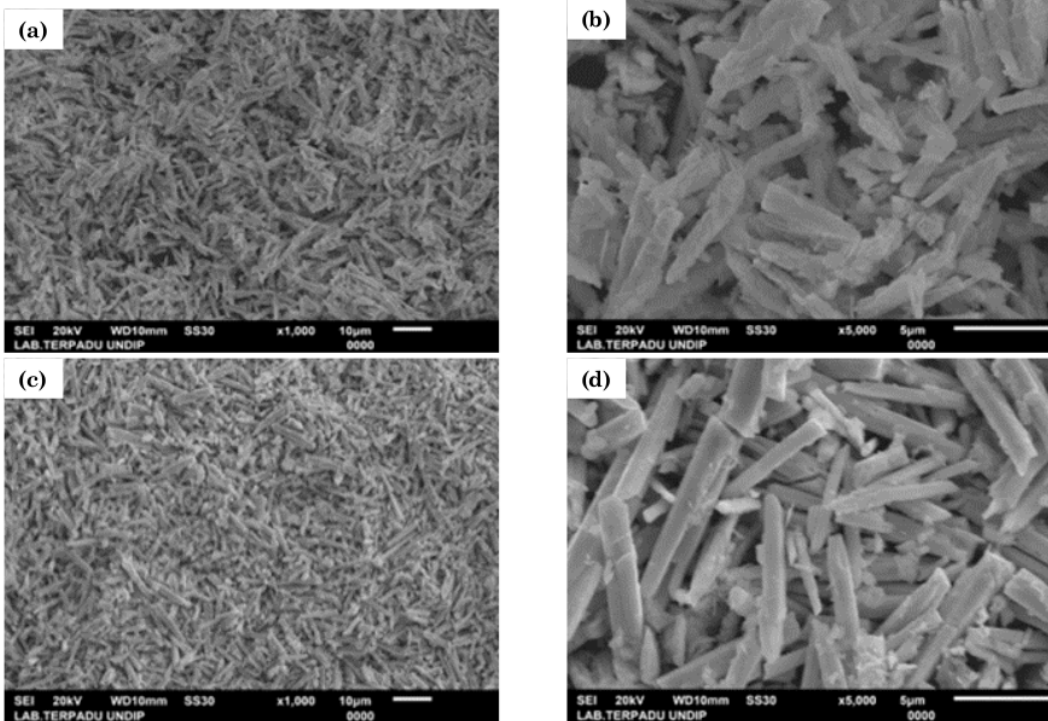


Figure 4. SEM images of bismuth oxide: (a) Bi12 at 1000x magnification; (b) Bi12 at 5000x magnification; (c) Bi16 at 1000x magnification; (d) Bi16 at 5000x magnification.

μm and 0.9–2.1 μm, respectively. Bi12 has a smaller particle size than Bi16. Particles with small sizes will have large surface areas, thus increasing the rate of photodegradation. This is consistent with the photocatalytic activity discussed in section 3.3 shown by Bi12 which has the highest activity among the samples.

The product bandgap with the highest photocatalyst activity, namely Bi₂O₃ (12 h) was determined using UV-DRS. The UV-DRS spectra of bismuth oxide with a hydrothermal time of 12 h are shown in Figure 5. The bandgap values were obtained using the Tauc Plot calculation method [30]. This was done by making a graph of the relationship between $h\nu$ and $(h\nu\alpha)^{1/2}$ and drawing a straight line that runs across the inflection point on the curve and intersects the energy axis. The x-axis represents the band (E_g) in units of eV, while the y-axis represents the value of $(h\nu\alpha)^{1/2}$. The Tauc equation is expressed in Equation (5).

$$(h\nu\alpha)^{1/2} = A(h\nu - E_g) \quad (5)$$

where, h = Planck's constant, ν = frequency, E_g = bandgap, and A = constant of proportionality. The exponential value n shows the nature of the sample transition.

The gridline that intersects the x-axis pointed out the band gap value of the bismuth oxide synthesized. Bi12 has two band gap values, namely 2.74 eV and 2.87 eV. The two band gap

values materialized because the synthesized Bi₂O₃ had a mixed crystal system, making it possible to have more than one band gap value. The bandgap values obtained also reinforced the fact that the bismuth oxide synthesized with a hydrothermal time of 12 h was mostly dominated by the monoclinic crystal system (α -Bi₂O₃) compared to other crystal structures in the sample since α -Bi₂O₃ has a bandgap value of 2.85 eV [11]. On the other hand, the bandgap value of 2.74 eV further confirmed the presence of a cubic crystal system (γ -Bi₂O₃) in the synthesized bismuth oxide, where γ -Bi₂O₃ has a bandgap of 2.783 eV [31].

Furthermore, the electrical conductivity of bismuth oxide prepared by hydrothermal method can be seen in Table 2. Compared to bismuth oxide prepared with other methods, hydrothermal method results in bismuth oxide with the highest electrical conductivity. The high electrical conductivity presumably related with the morphology of the particles in which French-fries like shape is like wire [32,33].

3.3 Photocatalytic Activity of Bismuth Oxide

The synthesized bismuth oxides were tested for their photocatalytic activity in degrading the methyl orange solution under ultraviolet light using a UV-A lamp (352 nm) with a time of 2-10 h. The MO solution used for the photocatalytic test had a concentration of 5 ppm and a volume of 33.0 mL for each variation of the test time. The absorbance of the degraded methyl orange solution was measured using a UV-Vis spectrophotometer at an optimum wavelength of 463 nm. The graph of the percentage degradation of methyl orange against photocatalytic time on the percentage is shown in Figure 6.

Figure 6 shows that at the 2 h test time, the five samples had relatively the same percentages of degradation. At the test time of 4 and 6 h, significant differences in the percentages of degradation were starting to be seen in the five samples. At 4 h, the percentage degradation of MO by Bi8, Bi10, Bi12, Bi14, and Bi16 were 11.25%, 15.31%, 22.50%, 23.13%, and 10.94%, respectively. Whereas at the 6 h time, the percentage of MO degradation by Bi8, Bi10, Bi12, Bi14, and Bi16 were 21.56%, 17.19%, 27.19%,

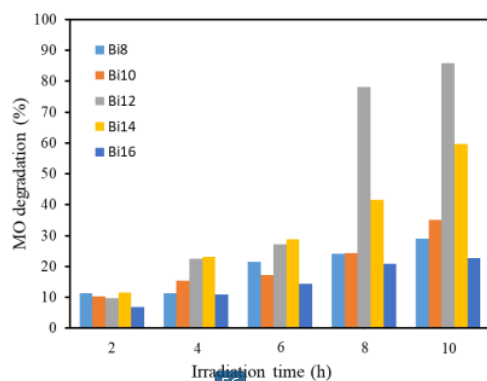


Figure 6. Graph of the effect of photocatalyst time on the percentage of MO degradation.

Table 2. Electrical properties of bismuth oxide prepared by different methods.

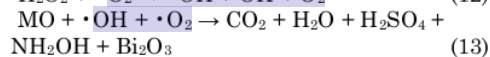
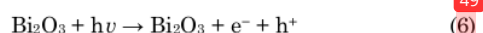
No	Products	Electrical conductivity (S.m ⁻¹)
1	Bi ₂ O ₃ prepared by hydrothermal	1.565×10 ⁻⁵
2	Bi ₂ O ₃ prepared by sol gel [6]	5.011×10 ⁻⁷
3	Bi ₂ O ₃ prepared by solution combustion [34]	1.367×10 ⁻⁷
4	Pure bismuth oxide from Sigma Aldrich	1.556×10 ⁻⁷

28.75%, and 14.38%, respectively. The percentage of MO degradation by bismuth oxide exhibited very significant differences at the test times of 8 h and 10 h. In the photocatalyst test with a time of 8 h, the percentage of MO degradation by Bi8, Bi10, Bi12, Bi14, Bi16 were 24.06%, 24.38%, 78.13%, 41.56%, and 20.94%, respectively. Whereas in the photocatalyst test with a time of 10 h, the percentage of MO degradation by Bi8, Bi10, Bi12, Bi14, Bi16 were 29.06%, 35.00%, 85.94%, 59.69%, and 22.81%.

Based on the above data, the bismuth oxide which had the best photocatalytic activity was Bi12. Additionally, the photocatalytic test shows presented that optimal photodegradation can be obtained at 10 h photocatalytic time. Bi12 has the highest photocatalytic activity in degrading the methyl orange compound. This is in accordance with the FTIR data, where the bismuth oxide with the heating time of 12 h had the smallest absorption intensity of Bi-OH groups, meaning that the amount of Bi(OH)₃ dehydrated into Bi₂O₃ was largest among other heating time variations. Also, XRD analysis data showed that Bi12 was predominated by the monoclinic crystal system (α -Bi₂O₃) compared to other crystal systems. Furthermore, the cubic crystal system (γ -Bi₂O₃) in Bi12 was highest in content when compared to Bi₂O₃ with other hydrothermal times. The data is also consistent with the UV-DRS analysis data in which bismuth oxide with the hydrothermal time of 12 h had two band gap values, namely 2.87 eV, which corresponds to the bandgap of α -Bi₂O₃ [11] and 2.74 eV, which corresponds to the bandgap of γ -Bi₂O₃ [31]. The number of cubic crystal systems (γ -Bi₂O₃) which was highest

compared to the other time variations is also in [46] with the fact that Bi12 had the highest photocatalytic activity. This can be attributed to the fact that γ -Bi₂O₃ has a smaller bandgap than α -Bi₂O₃. The smaller the band gap value, the better the photocatalytic process will be, because the electrons in the semiconductor will be more easily excited into the conduction band to form electron-hole pairs [35]. In addition, SEM results (see Figure 4) show that Bi12 has a relatively small particle size. When a material has a small particle size, then the material has a large surface area.

The mechanism of degradation of methyl orange (MO) by bismuth oxide that occurred is as follows:



When the Bi₂O₃ semiconductor was illuminated by light, the electrons in the valence band would be excited towards the conduction band. This resulted in the formation of hole and electron pairs. Furthermore, $\cdot\text{OH}$ originating from the water or water molecules adsorbed onto the surface of the Bi₂O₃ semiconductor would react with [45] holes to form $\cdot\text{OH}$ and H^+ . The electrons in the conduction band would react with atmospheric O_2 to form $\cdot\text{O}_2$. The $\cdot\text{O}_2$ could also

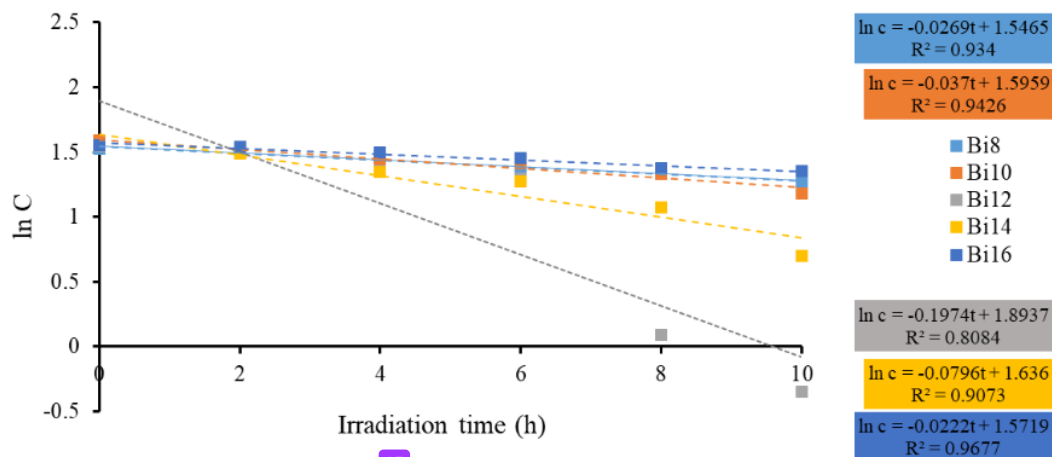


Figure 7. First-order reaction graph for the degradation of methyl orange.

serve as a source in the formation of $\cdot\text{OH}$ through a protonation process. The $\cdot\text{OH}$ formed would then oxidize the methyl orange molecule into simpler compounds, namely CO_2 and H_2O [36].

Dye degradation reactions generally follow the first-order reaction kinetics which is stated as follows [37]:

$$\ln C_t = \ln C_0 - kt \quad (14)$$

where, k = constant rate of the first order reaction (s^{-1}), C_0 = initial concentration of methyl orange solution (ppm), and C_t = concentration of methyl orange solution (ppm) at time t . The first-order reaction graph for the degradation of methyl orange dye by bismuth oxide synthesized is shown in Figure 7. The slope value of the line equation obtained in the graph implied the rate constant value of the methyl orange degradation reaction. Based on the graph, the rate constants of the degradation reaction of methyl orange by Bi8, Bi10, Bi12, Bi14 and Bi16 were $8.38 \times 10^{-6} \text{ s}^{-1}$, $1.05 \times 10^{-5} \text{ s}^{-1}$, $6.92 \times 10^{-5} \text{ s}^{-1}$, $2.56 \times 10^{-5} \text{ s}^{-1}$ and $6.86 \times 10^{-6} \text{ s}^{-1}$. It can be seen that Bi12 has the largest rate constant for the degradation of methyl orange. This corresponds to the photocatalytic test results where Bi12 has the highest photocatalytic activity in the degradation of methyl orange.

4. Conclusion

The synthesis of bismuth oxide using the hydrothermal method has been successfully carried out and it was proven that variations in hydrothermal time affect the characteristics of the resulting product. The best product characteristics were obtained in the bismuth oxide with the hydrothermal time of 12 h, shown by the results of FTIR analysis, where the Bi-O-Bi absorption in the sample was highest in intensity compared to others, and the SEM analysis results, which showed that the sample had a relatively small particle size (5.3–8.3 μm). Moreover, compared to bismuth oxide synthesized using different method, hydrothermal method resulted in bismuth oxide with higher electrical conductivity. Based on the results of the photocatalytic activity test, the bismuth oxide with the hydrothermal time of 12 h had the largest degradation reaction rate constant compared to other samples, which was $6.92 \times 10^{-5} \text{ s}^{-1}$.

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